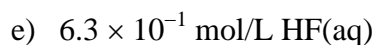
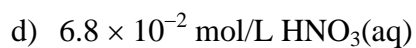
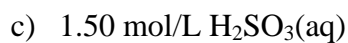
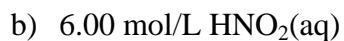
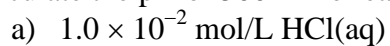


### **WS 5 - Strengths of Acids**

1. Calculate the pH of 500 mL of each of the following acids:



2. A 0.80 mol/L solution of an unknown acid, HX(aq), has a pH of 3.75. Calculate the K<sub>a</sub>, the acid ionization constant.

3. Calculate the pH of a solution containing 0.25 mol/L of an acid with an acid ionization constant (K<sub>a</sub>) of  $3.2 \times 10^{-6}$  mol/L.