



Remember this stuff from Sci 10

What is specific heat capacity?

The amount of **energy** (in J) needed to change **1g** of a substance by **1°C**

Specific heat capacity has the units **J/g °C**

The specific heat capacity of many elements/compounds is found on pg 3 of your data booklet

Which metal , copper or aluminum would get hotter if they both sat in the sun for the same period of time??

Q=mcΔt

How can we use the specific heat capacity of something to figure out how much energy it has gained or lost?

The formula $Q=mc\Delta t$ will help us do this

$Q =$ **amount of energy (in J)**

$m =$ **mass of substance (in g)**

$c =$ **specific heat capacity (in J/g °C)**

$\Delta t =$ **the temperature change in the substance in (°C)**

Determine the change in thermal energy when 115 mL of water is heated from 19.6 °C to 98.8 °C.