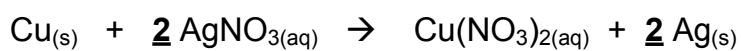


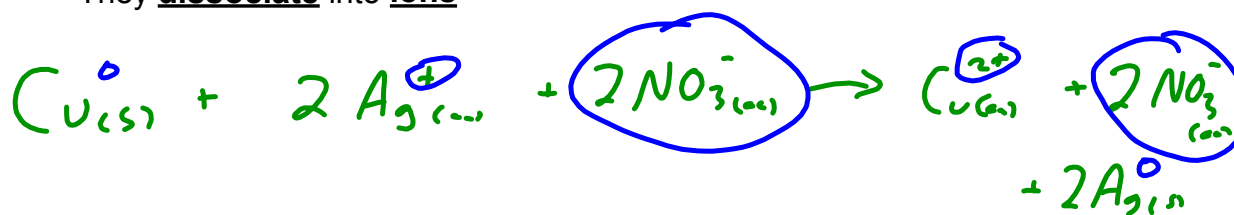
Lesson 7: Net Ionic Reactions pgs 281-283

Examine the following balanced reaction



What happens to aqueous ionic compounds?

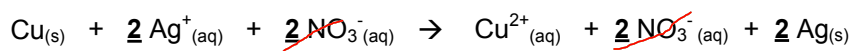
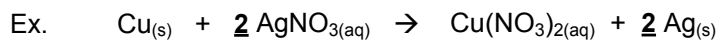
They **dissociate** into **ions**



- **Spectator Ion** – an ion that is not involved in a reaction (i.e. does not change from reactants to products)

- **Net ionic reaction** – a reaction that only shows the elements/ions that are involved in a reaction
(Spectator ions are not shown)

HCl can
H⁺ Cl⁻



- The net ionic reaction for the above reaction is:



* only split aqueous ionic compounds/acids

Example

Determine the net ionic reaction for the following reaction

Solid zinc is added to aqueous copper (II) sulfate.

