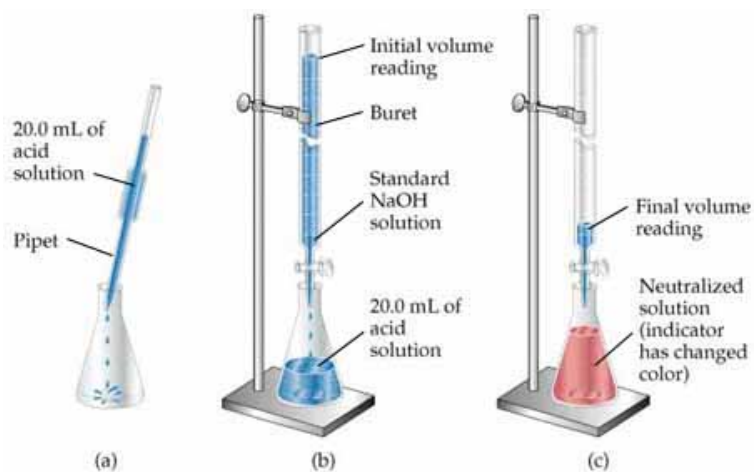
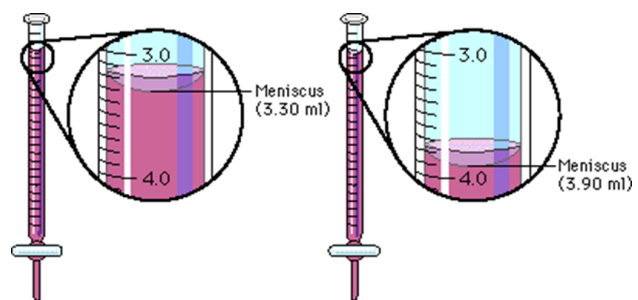
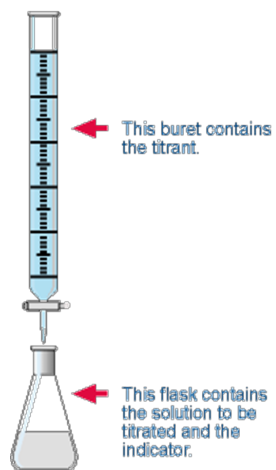


Lesson 6: Titration Analysis

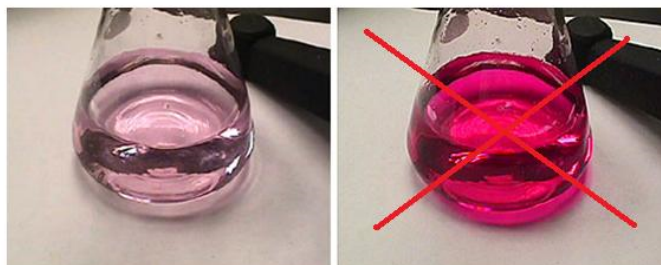
- **Titration** – process of carefully measuring and controlling the addition of a solution called the **titrant** from a **burette** into a measured, fixed volume of another solution called the **sample**
- **Endpoint** – a sudden observable change that indicates we have reached the equivalence point
- At the endpoint, the titration is stopped and the **volume** of the titrant added is determined





Important Info

- When doing a titration, you need to repeat the procedure more than once
- You need to obtain **3 values** for the volume of the titrant that are within 0.20mL of each other.
- The average is then used in the stoichiometry calculation
- Indicators are often used to signal the endpoint of a titration



Good Endpoint

Bad Endpoint (Overly Titrated)

